

grounds that it leads in a natural way to the intermolecular exchange observed at higher temperatures.

The chemistry of these new compounds as well as some related materials<sup>22</sup> is under active investigation. Preliminary results indicate that they possess unusual coordination preferences (SbF<sub>6</sub><sup>-</sup> over ethanol, for instance), that they regioselectively catalyze cyclopentadiene polymerization in high yield,<sup>23</sup> and that they react with metal anions resulting in both dimer formation and nitrosyl transfer. These results will be reported in due course.

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(22) For instance, use of the weaker donor ligand acetonitrile (AN) gives analogous chemistry when ANW(CO)<sub>5</sub> is combined with NO<sup>+</sup>BF<sub>4</sub><sup>-</sup>, although some disproportionation to give *mer,cis*-AN<sub>2</sub>(CO)<sub>3</sub>W(NO)<sup>+</sup>BF<sub>4</sub><sup>-</sup> occurs. With NO<sup>+</sup>SbF<sub>6</sub><sup>-</sup>, only disproportionation is observed, so acetonitrile apparently is not a strong enough donor to stabilize the W(μ-F)SbF<sub>5</sub> interaction.

(23) Farahi, J.; Hersh, W. H., unpublished results.

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**Supplementary Material Available:** Tables of crystallographic data for **1a** including details of the solution, positional parameters, temperature factors, bond lengths and angles, least-squares planes, and observed and calculated structure factors and theoretical NMR line shapes and rates for **1b** (25 pages). Ordering information is given on any current masthead page.

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## Additions and Corrections

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**Mechanism of the Reaction of Gas-Phase Iron Ions with Neutral Olefins** [*J. Am. Chem. Soc.* **1984**, *106*, 4307–4316]. D. A. PEAKE, M. L. GROSS,\* and D. P. RIDGE\*

Page 4314: In the paragraph titled Study of Products from Decomposition of FeC<sub>8</sub>H<sub>16</sub><sup>+</sup>, we refer to the product from reaction 17 as FeC<sub>6</sub>H<sub>12</sub><sup>+</sup>, a 1-hexene complex. The Fe(1-octene)<sup>+</sup> gives a FeC<sub>5</sub>H<sub>10</sub><sup>+</sup> fragment which is identical with Fe(1-pentene)<sup>+</sup>. We apologize for any confusion this may have caused.

**Nucleophilic Reactions in Solutions of Nonmicellized Hydrophobic Ammonium Ions** [*J. Am. Chem. Soc.* **1984**, *106*, 7178]. GIRMA BIRESAW, CLIFFORD A. BUNTON,\* CLIFFORD QUAN, and ZHONG-YUN YANG

Page 7182: Table VI footnotes are as follows. <sup>a</sup>Values are 10<sup>3</sup>kΨ, s<sup>-1</sup> at 25.0 °C. <sup>b</sup>In H<sub>2</sub>O:MeCN 70:30 (v/v), 0.01 M NaOH and 0.1 M **1e**. <sup>c</sup>In 2.2 vol % MeCN, 10<sup>-3</sup> M NaOH, 6 × 10<sup>-4</sup> M **1e** and 6 × 10<sup>-3</sup> M **1c**.

**Gas Phase <sup>1</sup>H NMR Spectra and Nitrogen Inversion of Trimethylene Imine** [*J. Am. Chem. Soc.* **1984**, *106*, 7638–7639]. B. RI. FRIEDMAN, J. PAUL CHAUVEL, JR., and NANCY S. TRUE\*

The temperature dependent exchange broadened <sup>1</sup>H NMR spectra reported in this communication result from the *cis* = *trans* nitrogen inversion process for α-methylaziridine. They were incorrectly attributed to nitrogen inversion in the four-membered-ring azetidone. The authors thank F. A. L. Anet of UCLA for bringing this fact to their attention as well as for providing the right assignment of the spectra.